

Faculty of Chemical & Natural Resources Engineering

Universiti Malaysia Pahang

SCIENCE & ENGINEERING MATERIAL

**Quiz Chapter 1** 

## Question

For a 100 g sample of nickel aluminide (NiAl<sub>3</sub>) compound, calculate

- a) Mol of nickel aluminide
- b) Quantity of nickel aluminide molecules, nickel atoms, and aluminium atoms in the sample

Avogadro's number =  $6.023 \times 10^{23}$  atoms/mol Molecular weight of nickel = 58.71 g/mol Molecular weight of aluminium = 26.98 g/mol